BRAINWARE UNIVERSITY

Term End Examination 2018 - 19

Programme– Bachelor of Pharmacy

Course Name – Pharmaceutical Inorganic Chemistry

Course Code - BP104T

(Semester – 1)

Time allotted: 3 Hours **Full Marks: 75**

[The figure in the margin indicates full marks. Candidates are required to give their answers in their own words as far as practicable.]

Group –A

(Multiple Choice Type Question) 20 x 1 = 20

1. *Choose the correct alternative from the following*

(i) First I.P published in the year

a. 1945 b. 1950
c. 1947 d. 1955

(ii) Latest I.P published in the year

a. 2007 b.2014
c. 2017 d. 2018

(iii) pH of 0.1N HCL

a. 1 b. 0
c. 7 d. 14

(iv) The combination of aluminum and magnesium salts as antacid are used to

a. To increase the activity of each antacid
b. Balance the constipation and laxative effect
c. Increase the absorption of each antacid
d. None of these

(v) Isotonic solution of sodium chloride is

a. 0.9%W/V b. 1.5%W/V
c. 0.4%W/V d. 2.0%W/V

(vi) Ammonium Chloride used as

a. Systemic Acidifiers
b. Expectorant
c. Diuretic
d. All of above

(vii) pH of Blood is

a. 7.4
b. 7.8
c. 4.5
d. 7.0

(viii) Muriatic acid is known as

a. HCl
b. H_2SO_4
c. HNO3
d. None of these

(ix) HCl content in dilute HCl is

a. 2.0%w/v
b. 10%w/v
c. 5.0%w/v
d. 15%w/v

(x) Which is used as gastric antacid

a. $Mg(OH)_2$
b. $NaHCO_3$
c. $MgSO_4$
d. CH_3COONa

(xi) The indicator used during Assay of $KMnO_4$

a. Self-Indicator
b. Mucilage of Starch
c. Methyl Orange
d. None of these

(xii) The use of Sodium potassium Tartarate is

a. Saline Cathertic
b. Urinary Alkaliser
c. Mild Laxative
d. All of the above

(xiii) In I.P Refrigerated Temp. is

a. 1 to 10°
b. 0 to 5°C
c. 2 to 8°C
d. 2 to 10°C

(xiv) Measurement of Radioactivity by

a. Ionization Chamber
b. G.M Counter
c. Scintillation counter
d. All of above

(xv) Heavy metals are those whose density greater than

a. 20
b. 10
c. 5
d. 30

(xvi) Limit of Arsenic in drinking water as per I.P is less than

a. **5** ppm b. 01 ppm

c. 10 ppm d. 02 ppm

(xvii) Essential to regulate Acid Base Balance

a. Sulphate b. Bicarbonate

c. Chloride d. Sodium

(xviii) Which of the following is a systemic antacid

a. $CaCO_3$ b. KMno4

c. $NaHCO_3$ d. None of these

(xix) Thioglycolic acid is used for the limit test of

a. Iron b. Lead

c. Chloride d. Heavy metal

(xx) pH range of Urine

a. 4-6 b. 7-10

c. 6-8 d. 4.5-8

Group – B

(Short Answer Type Questions) 7 x5 = 35

Answer any *seven* from the following

2.	What are anti caries agents? Explain the role of Fluoride ions	5
3.	Define Antidotes. Classify Antidotes according to their examples.	5
4.	Why antacids are given in combination always? Explain with some marketed preparations.	5
5.	Write a note on Dentifrices and Desensitizing Agents.	5
6.	Define with examples a) Expectorant b) Emetic c) Astringents	5
7	Derive Henderson –Hesselbatch equation for Acidic Buffer solution	5
8	Write the composition of ORS as per WHO and I.P with the uses of the ingredients.	5
9	Define Isotonic, Hypotonic and Hypertonic solutions with examples.	5
10	Why Nitric Acid is used for the limit test of Chloride?	5

Group – C

(Long Answer Type Questions) 2x 10 = 20

Answer any *two* from the following

11.	(a)	Write briefly the modern concept of Acids and Bases with examples.	4
	(b)	Define pH , pOH and Buffer solution with examples.	4
	(c)	Write the role of Buffer solution in Pharmacy.	2
12.	(a)	Define Limit Test.	1
	(b)	Why Limit test is done?	2
	(c)	Write the Principle, reaction and procedure of the Limit test of Arsenic I.P. Describe the apparatus involved.	7
13.	(a)	Discuss all the major physiological ions with their functions	6
	(b)	Write a detail note on physiological buffers.	4
